

Investigation: (Do 1-9, skip 5b and 5c)

1.a. Use large test tube, make a prediction

H ₂ O amount	Vinegar amount	Reaction time (don't wait until all dissolve, record when first one dissolves, observe other two)
0 mL	20 mL	
10 mL	10 mL	
15 mL	5 mL	

1.b. (answer)

1.c. compare to prediction

2.a. Use well plate (each well holds 3.5 mL)

H ₂ O amount	Vinegar amount	Reaction time (don't wait until all dissolve, record when first one dissolves, observe other two)
0 mL	3 mL	
1 mL	2 mL	
2 mL	1 mL	

2.c. compare the results to 1.b. What was the difference in using Large test tube vs smaller well plate, even though same experiment?

Name _____ Period _____

Chemistry 1 Chapter 6 Activity 6 Investigate page 477 (attach to lab)

Write a balanced equation for the following steps to Chapter 6 Activity 6, identify the gas produced and how you could identify it using a splint (be specific).

Investigate number 1 Magnesium and vinegar (acetic acid $\text{HC}_2\text{H}_3\text{O}_2$)

Balanced equation _____

What Gas was Produced? _____

How can you identify the gas? (with the Splint) _____

Investigate number 3 HCl and Zn

Balanced equation _____

What Gas was Produced? _____

How can you identify the gas? (with the Splint) _____

Investigate number 4 Hydrogen peroxide (Formula on page 478)

Balanced equation _____

What Gas was Produced? _____

How can you identify the gas? (with the Splint) _____

Investigate number 7 water and antacid tablet (sodium bicarbonate NaHCO_3)
Don't need to balance equation, just identify gas (look at the formula)

What Gas was Produced _____

How can you identify the gas? _____