

Chapter 6 Activity 4 Investigate page 459

	Silver Nitrate AgNO_3	Copper(II) sulfate CuSO_4	Magnesium sulfate MgSO_4	Sodium Hydroxide NaOH
Potassium carbonate K_2CO_3	Results 1)After mixing	Results 2)After mixing	Results 3)After mixing	Results 4)After mixing
Sodium hydroxide NaOH	Results 5)After mixing	Results 6)After mixing	Results 7)After mixing	Results 8)After mixing
Potassium Iodide KI	Results 9)After mixing	Results 10)After mixing	Results 11)After mixing	Results 12)After mixing
Iron(III) chloride FeCl_3	Results 13)After mixing	Results 14)After mixing	Results 15)After mixing	Results 16)After mixing

11c. Write the balanced equation for each reaction using the same number as in the chart above. If no reaction, write "no reaction". For each of the products, identify which is the precipitate (s) and which is aqueous (aq) using the chart on page 460.

1. _____

2. _____

3. _____

4. _____

5. _____

6. _____

7. _____

8. _____

9. _____

10. _____

11. _____

12. _____

13. _____

14. _____

15. _____

16. _____