

Understanding Ionic, Molecular and Acidic compounds

1. HF
2. $\text{HC}_2\text{H}_3\text{O}_2$
3. LiI
4. ZnO
5. H_2O
6. KOH
7. $\text{NaC}_2\text{H}_3\text{O}_2$
8. $\text{Ba}(\text{OH})_2$
9. FeCl_3
10. HgO
11. FePO_4
12. CO
13. $\text{Ba}(\text{BrO}_3)_2$
14. HgOH
15. HgBr_2
16. HNO_3
17. NH_4Cl
18. $\text{Al}_2(\text{SO}_4)_3$
19. NH_4NO_3
20. NI_3
21. CaSO_4
22. NaHCO_3
23. $(\text{NH}_4)_3\text{PO}_4$
24. NH_4OH
25. PbSO_4
26. Cu_2S
27. $\text{Sn}(\text{HCO}_3)_4$
28. $\text{Zn}(\text{NO}_2)_2$
29. Hg_2SO_4
30. Al_2O_3
31. $\text{Co}_2(\text{SO}_3)_3$
32. SnCrO_4
33. $\text{Pb}(\text{OH})_2$

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34. KF

35. HCl

36. H_3PO_4

37. K_2O

38. NaCl

39. CuCO_3

40. H_2SO_4

41. H_2CO_3

Write the formula for the following compounds

42. Lithium sulfate

43. phosphorus trichloride

44. zinc hydroxide

45. ammonium bisulfate

46. diphosphorus pentoxide

47. bromic acid

48. lead(IV) chloride

49. mercury (I) sulfate

50. magnesium sulfite

51. zinc acetate

52. calcium hydride

53. silicon dioxide

54. cesium carbonate

55. nitric acid

56. silver(I) oxide

57. tin (II) bromide

58. dinitrogen pentoxide

59. potassium permanganate

60. ammonium bisulfate

61. lead (IV) chloride

62. ammonium hydroxide

63. antimony pentasulfide

64. copper (II) nitrite