

## Formulas and Nomenclature

Name the following compounds according to the rules for ionic, molecular, and acidic compounds.

### Ionic Compounds

1. KOH
2. HgOH
3. KCl
4. FeCl<sub>3</sub>
5. NH<sub>4</sub>OH
6. Cu<sub>2</sub>O
7. Al<sub>2</sub>(SO<sub>4</sub>)<sub>3</sub>
8. NaOH
9. Pb(OH)<sub>2</sub>
10. NH<sub>4</sub>NO<sub>3</sub>
11. NaHCO<sub>3</sub>
12. HgO
13. Zn(NO<sub>2</sub>)<sub>2</sub>
14. CsOH
15. Li<sub>2</sub>O
16. Ca(OH)<sub>2</sub>
17. CaBr<sub>2</sub>
18. Fe<sub>2</sub>O<sub>3</sub>
19. FeCO<sub>3</sub>
20. Ba(BrO<sub>3</sub>)<sub>2</sub>
21. Al(OH)<sub>3</sub>
22. NaC<sub>2</sub>H<sub>3</sub>O<sub>2</sub>
23. Na<sub>2</sub>SO<sub>3</sub>
24. LiH
25. MgBr<sub>2</sub>
26. SnBr<sub>2</sub>
27. NH<sub>4</sub>F
28. KHCO<sub>3</sub>
29. K<sub>2</sub>O
30. ZnO
31. NaClO
32. SrS
33. Al(BrO<sub>3</sub>)<sub>3</sub>
34. Pd(CN)<sub>2</sub>
35. ZnSiO<sub>3</sub>
36. Mg(C<sub>2</sub>H<sub>3</sub>O<sub>2</sub>)<sub>2</sub>

37.  $\text{Ca}(\text{MnO}_4)_2$

38.  $\text{Be}(\text{NO}_3)_2$

39.  $\text{RaBr}_2$

40.  $\text{NaMnO}_4$

41.  $\text{PbI}_2$

42.  $\text{CaS}$

43.  $\text{KClO}_4$

44.  $\text{HgBr}_2$

45.  $\text{CoSi}$

46.  $\text{CuSO}_3$

47.  $\text{FePO}_4$

48.  $\text{PbTe}$

49.  $\text{HgNO}_3$

50.  $\text{K}_2\text{SiO}_3$

51.  $\text{AgC}_2\text{H}_3\text{O}_2$

52.  $\text{Zn}_3(\text{PO}_4)_2$

53.  $\text{Ag}_2\text{S}$

54.  $\text{Cd}(\text{HCO}_3)_2$

55.  $\text{ZnF}_2$

Molecular Compounds

56.  $\text{N}_2\text{O}_5$

57.  $\text{CO}_2$

58.  $\text{SO}_3$

59.  $\text{CO}$

60.  $\text{N}_2\text{O}$

61.  $\text{AsCl}_5$

62.  $\text{SbF}_3$

63.  $\text{Bi}_2\text{Te}_3$

64.  $\text{P}_3\text{N}_5$

65.  $\text{TeI}_4$

Acidic Compounds

66.  $\text{HCl}$

67.  $\text{HNO}_3$

68.  $\text{HF}$

69.  $\text{H}_3\text{PO}_4$

70.  $\text{H}_2\text{SO}_4$

71.  $\text{H}_2\text{CO}_3$